

Lability and mobility effects on mixtures of ligands under steady-state conditions.

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Abstract

Analytical solutions for the steady-state flux arriving at an active surface from a mixture (in which one active species reacts with non-active ligands in the medium) can be helpful in a variety of problems: voltammetric techniques, heterogeneous processes in reactors, toxic or nutrient uptake, techniques of diffusive gradients in thin films (DGT), etc. Under any geometry that sustains steady-state, a convenient combination of the reaction-diffusion equations leads to a simpler formulation of the problem for arbitrary diffusivities of the species and arbitrary rate constants of the first order conversion between the active species and the non-active species. The resulting problem can be characterised in terms of a list of dimensionless parameters involving the kinetic and mobility properties of each species. A lability degree for each 1:1 complex in terms of the surface concentrations leads to: i) a lability criterion specific for each complex in the mixture and ii) the assessment of the relative contribution of each complex to the resulting flux. Semi-infinite spherical diffusion (as in the Gel Integrated MicroElectrode, GIME, biouptake modelling of micro-organisms, etc.) is specifically considered and some consequences of its full analytical solutions are discussed.

Keywords: reaction-diffusion, steady state, macromolecular complexation, GIME, uptake

Introduction

Diffusion coupled with chemical reaction in the solution appears in a variety of problems in different scientific areas such as:

i) Cellular nutrition ^[1-4] and biouptake of pollutants by micro-organisms ^[5-7]. In many cases, the molecules of nutrients or toxins can be complexed by ligands present in the medium while the diffusive transport occurs. The bioavailability of the substance of interest will depend not only on the nature of the active surface (the cellular membrane in this case) and the dynamics of diffusion, but also on the extension and the kinetics of the complexation processes in the medium (i.e. its kinetic speciation ^[8,9]).

ii) Heterogeneous catalysis and electrode reactions. The problem is also relevant to Chemical Engineering ^[10] because the efficiency of the heterogeneous reaction is affected by diffusion and reaction in the solution phase.

iii) Analytical instrumental measurements. The kinetic contribution of dissociating complexes to the measured currents in polarography was already recognised by Koutecky and other members of the Czechoslovak school ^[11], became an important topic ^[12-16] and has also received recent attention ^[17-19]. New analytical techniques, such as DGT ^[20-22], gel integrated voltammetric microelectrodes (GIME)^[23] and Permeation Liquid Membrane ^[24] also have to tackle reaction coupled to diffusion in interpreting the measured signals.

In all of these systems, the responses due to the arrival of the active substance at the surface depend not only on the equilibrium values of its complexation with the ligands (thermodynamic speciation), but also on the particular values of the various association

and dissociation rate constants, i.e. on the dynamic speciation. Two limiting cases are easily recognised: the inert limit (no dissociation of the complexes) and the fully labile limit (the equilibrium relationship holds at any point of the solution). Thus, the term lability can be understood as a continuous property (between the aforementioned limits, passing through different amounts of partial lability) measuring the contribution of the dissociation of the complexes to the response flux^[25;26]. In this context, lability depends not only on solution properties (values of diffusion coefficients, complexation and rate constants, etc.) but also on the time scale of the experiment, the geometry of the surface, the thickness of the diffusion domain, etc.

The aim of this work is to study steady-state fluxes arriving at an active surface due to diffusion and reaction of an active species and non-active ligands, taking into account all the dynamic properties of the mixture, including different mobilities and no restriction for the values of the association/dissociation rate constants. We extend the procedure of Turner and Whitfield^[12] who developed a very general solution of multicomponent reaction, convection and diffusion (with all components sharing a common diffusion coefficient). By considering diffusion as the only transport phenomenon, we can deal with unequal mobilities (i.e. unequal diffusion coefficients) for the different complexes, so that macromolecular complexation –which is essential for understanding environmental problems^[27] - can be realistically modelled.

For simplicity, but without losing the generality of the above mentioned applications, we use the voltammetric terminology. The active surface is the electrode, which is the sink of the metal ion M (active species). In the solution, M can be complexed by h ligands to form h different complexes, none of which is (electro)active on the electrode

surface. In the given examples, we will restrict ourselves to fluxes in diffusion limited conditions (null concentration of M at the active surface). The extension of the treatment here expound to the case of a fixed arbitrary value for the concentration of the active species at the active surface (due to phenomena such as Michaelis-Menten internalisation or a Nernstian relationship between species at the surface) is straightforward and is not detailed here for simplicity reasons.

In the first section of this work, we lay down the re-formulated dimensionless equations that should be solved under any specified geometry. In section 2, we introduce the concept of degree of lability for each complex, which allows the formulation of lability criterion for each complex. In section 3, we concentrate on spherical geometry with semi-infinite diffusion, presenting closed expressions for a mixture of two ligands. Derivation of expressions for one dimensional geometries are included in Appendix A, while the symbols used are gathered in Appendix B.

The results of this article can be applied to discrete mixtures of ligands, either arising from the mixing of a number of different real species or from the consideration of sites in macromolecules as formal species which can have different affinities (e.g. if different functional groups are involved in different kinds of sites) but have the same mobility.

1.- General formulation for any geometry

1.1.- Setting of the problem

Let a multicomponent mixture consist of one active species (the free metal ion M in the adopted terminology) and h different ligands designated $L_1, L_2..L_h$, exhibiting h simultaneous complexation reactions of the type



where subscripts a and d indicate association and dissociation respectively. See Fig 1 for an outline of the kind of system considered. Assuming ideal dilute behaviour, the equilibrium relationship (fulfilled in the “bulk”, denoted with superscript *, either at a finite or infinite distance from the active surface) is:

$$K_i' \equiv K_i c_{L_i}^* = \frac{k_{a,i} c_{L_i}^*}{k_{d,i}} = \frac{c_{ML_i}^*}{c_M^*} \quad (2)$$

Assume that the ligand concentrations are large in comparison with that of M so that their value can be approximated as constants $c_{L_i}^*$ at any point of the solution (i.e. ligand excess conditions). The excess of ligand condition for a given ligand i is clearly fulfilled when $20c_{ML_i}^* < c_{L_i}^*$ implying $c_M^* < 1/(20K_i)$; but the ligand of excess approximation can yield good results for the flux even at lower ligand-to-metal ratios. When ligand excess conditions apply, the kinetics of interconversion of M and ML_i follow a pseudo first-order law and the problem becomes linear.

Assume a steady-state mass-conservation equation for each complex ML_i with diffusion as the only relevant transport phenomenon:

$$D_{ML_i} \nabla_{\text{dim}}^2 c_{ML_i} - k_{d,i} c_{ML_i} + k_{a,i} c_{L_i}^* c_M = 0 \quad (3)$$

where the (dimensional) Laplacian operator depends on the particular geometry considered. For the free metal ion, the mass-conservation equation reads:

$$D_M \nabla_{\text{dim}}^2 c_M + \sum_{i=1}^h k_{d,i} c_{ML_i} - \left(\sum_{i=1}^h k_{a,i} c_{L_i}^* \right) c_M = 0 \quad (4)$$

In order to work preferentially with dimensionless parameters we define:

$$\varepsilon_i \equiv \frac{D_{ML_i}}{D_M} \quad (5)$$

which is an index of the mobility of the particular complex i with respect to the metal ion (active species). For the association kinetics, we introduce

$$\kappa_{a,i} \equiv \frac{k_{a,i} c_{L_i}^* \ell^2}{D_M} \quad (6)$$

where ℓ stands for some characteristic length of the active area (such as the radius of the sphere or of the disc electrode) [28;29] or a characteristic length of the system (such as a diffusion layer thickness). $\kappa_{a,i}$ is a dimensionless association parameter [18;30], Damköhler number or Thiele modulus [10;31-34], comparing the diffusion efficiency of the metal and its association rate constant with a given ligand L_i . One noticeable fact is that $\kappa_{a,i}$ depends on the characteristic length: for decreasing values of ℓ , $\kappa_{a,i}$ decreases indicating that the role of diffusion increases with respect to the role of kinetics [18]. For a fixed metal and a fixed ℓ , the $\kappa_{a,i}$ -values are just proportional to the corresponding $k_{a,i}$ of the reaction of M with the different ligands.

For notation simplicity we also define

$$\kappa_0 \equiv \sum_{i=1}^h \kappa_{a,i} \quad (7)$$

Thus κ_0 indicates the added effect of all the ligands.

Similarly, for the dissociation process

$$\kappa_{d,i} \equiv \frac{k_{d,i} \ell^2}{D_{ML_i}} \quad (8)$$

which can be seen as the dimensionless dissociation parameter comparing, for a given complex, the effectiveness of the processes of its dissociation and its diffusion.

Notice that

$$\frac{\kappa_{a,i}}{\kappa_{d,i}} = \varepsilon_i K_i' \quad (9)$$

We also introduce the normalised concentrations:

$$c_i \equiv \frac{\varepsilon_i K_i' c_{ML_i}}{c_{ML_i}^*} \quad 0 < i \leq h \quad (10)$$

and

$$c_0 \equiv \frac{c_M}{c_M^*} \quad (11)$$

Then, eqns. (3) and (4) can be simply re-casted as

$$\nabla^2 c_i - \kappa_{d,i} c_i + \kappa_{a,i} c_0 = 0 \quad 0 < i \leq h \quad (12)$$

$$\nabla^2 c_0 + \sum_{i=1}^h \kappa_{d,i} c_i - \kappa_0 c_0 = 0 \quad (13)$$

where now the Laplacian operator is dimensionless.

We consider 2 response functions: i) the gradient of normalised concentration of metal $\text{grad}(c_0)$ at the electrode surface and ii) the normalised flux of metal ϕ , which is the ratio between the total flux measured with and without ligand ^[35;36] for the same total metal concentration

$$\phi \equiv \frac{\left[\int_{\substack{\text{active} \\ \text{Area}}} \text{grad}(c_0) dA \right]_{\text{with ligands}}}{\left[\int_{\substack{\text{active} \\ \text{Area}}} \text{grad}(c_0) dA \right]_{\text{without ligands}}} \quad (14)$$

which in voltammetry can be simply obtained as the ratio of currents with and without the ligand presence ^[15].

1.2.- Method of solution for any geometry

The set of $h+1$ differential equations (13) and (12) can be written in matrix form as:

$$\nabla^2 \vec{c} = \mathbf{A} \cdot \vec{c} \quad (15)$$

where

$$\mathbf{A} \equiv \begin{pmatrix} \kappa_0 & -\kappa_{d,1} & -\kappa_{d,2} & \cdots & -\kappa_{d,h} \\ -\kappa_{a,1} & \kappa_{d,1} & 0 & \cdots & 0 \\ -\kappa_{a,2} & 0 & \kappa_{d,2} & \cdots & 0 \\ \vdots & \vdots & \vdots & & \vdots \\ -\kappa_{a,h} & 0 & 0 & \cdots & \kappa_{d,h} \end{pmatrix} \quad (16)$$

We seek a combination of equations (12) and (13) so that they are re-written as

$$\nabla^2 f_j = n_j f_j \quad (17)$$

where f_j are some combinations of the concentrations c_i ; i.e. the goal is to uncouple the differential equations for some formal species concentrations f_j (also called “quasi-species” in the theory of macromolecular evolution ^[37]). In matrix form:

$$\nabla^2 \vec{f} = \mathbf{D} \cdot \vec{f} \quad (18)$$

where \mathbf{D} is a diagonal matrix. If \mathbf{A} can be rendered diagonal, there exists a matrix \mathbf{S} so that:

$$\mathbf{A} = \mathbf{S} \cdot \mathbf{D} \cdot \mathbf{S}^{-1} \quad (19)$$

Thus, the sought relationship between the concentrations can be expressed:

$$\bar{\mathbf{c}} = \mathbf{S} \cdot \bar{\mathbf{f}} \quad \bar{\mathbf{f}} = \mathbf{S}^{-1} \cdot \bar{\mathbf{c}} \quad (20)$$

or

$$c_i = \sum_{j=0}^h s_{ij} f_j \quad (21)$$

Finding the combination

The elements of the diagonal matrix \mathbf{D} are the eigenvalues of the matrix \mathbf{A} , which can be found as the zeros of the characteristic polynomial of \mathbf{A} . Following the usual procedure^[38], we have to solve:

$$\det(t \cdot \mathbf{I} - \mathbf{A}) = 0 \quad (22)$$

where \mathbf{I} is the identity matrix. In the determinant, we add all the rows to the first one taking into account the definition (7); we factor out t from the first row and factor out $(t - \kappa_{d,j-1})$ from each row j ($2 \leq j \leq h$), so that eqn. (22) can be expressed as

$$(t - \kappa_{d,1}) \cdot \dots \cdot (t - \kappa_{d,h}) \cdot t \cdot \begin{vmatrix} 1 & 1 & 1 & \dots & 1 \\ \frac{\kappa_{a,1}}{t - \kappa_{d,1}} & 1 & 0 & \dots & 0 \\ \frac{\kappa_{a,2}}{t - \kappa_{d,2}} & 0 & 1 & \dots & 0 \\ \vdots & \vdots & \vdots & \ddots & \vdots \\ \frac{\kappa_{a,h}}{t - \kappa_{d,h}} & 0 & 0 & \dots & 1 \end{vmatrix} = 0 \quad (23)$$

which, after subtracting all the rows from the first one, leads to

$$(t - \kappa_{d,1}) \cdot \dots \cdot (t - \kappa_{d,h}) \cdot t \cdot \left(1 + \sum_{i=1}^h \frac{\kappa_{a,i}}{\kappa_{d,i} - t} \right) = 0 \quad (24)$$

One solution of the previous eqn. is $t=0$, which can be labelled as the eigenvalue n_0 . The other h eigenvalues n_1, n_2, \dots, n_h required in eqn. (17) can be found from

$$1 + \sum_{i=1}^h \frac{\kappa_{a,i}}{\kappa_{d,i} - t} = 0 \quad (25)$$

which is a polynomial equation of degree h in t . We stress that n_1, n_2, \dots, n_h are function of $\kappa_{d,i}$ and $\kappa_{a,i}$, but totally independent of the geometry.

It is worth noticing that all the values n_1, n_2, \dots, n_h are positive, each one lying between two $\kappa_{d,i}$:

$$0 = n_0 < \kappa_{d,1} < n_1 < \kappa_{d,2} < n_2 \cdots < \kappa_{d,h} < n_h \quad (26)$$

This property can be deduced from considering the graphical solution of eqn. (25) in the following form

$$\sum_{i=1}^h \frac{t \kappa_{a,i}}{t - \kappa_{d,i}} = t \quad (27)$$

each n_j is the intersection between the bisectant of the first quadrant (r.h.s. in eqn. (27)) and a sum of hyperbola-like functions (l.h.s. in eqn. (27)), as seen in Fig 2 for the simple case of $h=3$. The knowledge that there is one eigenvalue n_j between two consecutive $\kappa_{d,i}$ can be exploited for accurately determining the n_j -values. On the other hand, since all the eigenvalues are different, matrix **A** can be diagonalized and all the construction is proved consistent.

In general, each n_j depends on the whole set of $\kappa_{a,i}$ -values and $\kappa_{d,i}$ -values, as indicated in eqn.(25). For a given formal species with concentration f_j , the parameter n_j can be physically understood, according to eqn. (17), as the non-dimensional kinetic constant

of disappearance of that formal species. Larger n_j -values correspond to formal species having thinner reaction layers.

Suppose that we know (e.g. numerically) the eigenvalues. Now, we turn to the problem of finding the coefficients s_{ij} in \mathbf{S} . As the columns in matrix \mathbf{S} are the eigenvectors^[38], we just need to solve the eigenvalue problem:

$$(n_j \cdot \mathbf{I} - \mathbf{A}) \bar{\mathbf{s}}_j = 0 \quad (28)$$

which is the system

$$\begin{pmatrix} n_j - \kappa_0 & \kappa_{d,1} & \kappa_{d,2} & \cdots & \kappa_{d,h} \\ \kappa_{a,1} & n_j - \kappa_{d,1} & 0 & \cdots & 0 \\ \kappa_{a,2} & 0 & n_j - \kappa_{d,2} & \cdots & 0 \\ \vdots & \vdots & \vdots & & \vdots \\ \kappa_{a,h} & 0 & 0 & \cdots & n_j - \kappa_{d,h} \end{pmatrix} \begin{pmatrix} s_{0,j} \\ s_{1,j} \\ s_{2,j} \\ \vdots \\ s_{h,j} \end{pmatrix} = \begin{pmatrix} 0 \\ 0 \\ 0 \\ \vdots \\ 0 \end{pmatrix} \quad (29)$$

As the previous system is undetermined, the first equation can be skipped and, as the eigenvectors must be non-null, we can choose

$$s_{0,j} = 1 \quad 0 \leq j \leq h \quad (30)$$

which leads to

$$c_0 = \sum_{j=0}^h f_j \quad (31)$$

and the other coefficients can be easily found from the remaining equations in (29):

$$s_{ij} = \frac{\kappa_{a,i}}{\kappa_{d,i} - n_j} \quad 0 < i \leq h \quad (32)$$

Computing the flux

We assume that M is the only active species on the active surface (electrode). Thus,

$$\text{grad}(c_i) = 0 \quad 0 < i \leq h \quad (33)$$

A useful result can be obtained by adding up

$$c_0 + \sum_{i=1}^h c_i = \sum_{j=0}^h f_j + \sum_{i=1}^h \left(\sum_{j=0}^h s_{ij} f_j \right) = \sum_{j=0}^h \left(1 + \sum_{i=1}^h \frac{\kappa_{a,i}}{\kappa_{d,i} - n_j} \right) f_j = \left(1 + \sum_{i=1}^h \varepsilon_i K_i' \right) f_0 \quad (34)$$

where eqn. (25) has been used. Thus,

$$\text{grad}(c_0) = \left(1 + \sum_{i=1}^h \varepsilon_i K_i' \right) \text{grad}(f_0) \quad (35)$$

which indicates that the flux of M can be easily computed once f_0 is known.

Outline of the solution procedure

Once the roots n_1, \dots, n_h in (25) are found, one must solve (17) with the appropriate boundary conditions:

BOUNDARY CONDITIONS FOR f_j AT THE ACTIVE SURFACE: On one hand, at the electrode surface, we have no flux given by (33) corresponding to c_i ($i > 0$). These conditions, referred to f_i via (21), become:

$$\text{grad}(c_i) = 0 = \sum_{j=0}^h s_{ij} \text{grad}(f_j) \quad 0 < i \leq h, \text{ active surface} \quad (36)$$

If we are under diffusion limited conditions, then $c_0=0$, which referred to f_j via (31) becomes

$$c_0 = 0 = \sum_{j=0}^h f_j \quad \text{active surface} \quad (37)$$

else some value should be provided for c_0 (directly or from complementary information such as Nernstian equilibrium). Notice that the boundary conditions to apply to each f_i at the active surface are implicit within the summations (36) and (37), i.e. the combination (changing from c_j to f_j) that simplifies the differential equations mixes up the boundary conditions everywhere (and thus at the active surface, here considered). So, in order to

prescribe the boundary conditions on a given f_j one needs, in general, some kind of inversion of the matrix \mathbf{S} (see eqn. (51) in ref. [12]).

BOUNDARY CONDITIONS FOR f_j IN THE BULK: On the other hand, more boundary conditions are needed, even for 1-D geometries. In some cases, fixed bulk conditions are prescribed, either at a finite or infinite distance. In these cases, eqn. (34) leads to the boundary condition

$$f_0^* = \frac{c_0^* + \sum_{i=1}^h c_i^*}{1 + \sum_{i=1}^h \varepsilon_i K_i'} = 1 \quad (38)$$

where definitions (10) and (11) have been applied.

For the other formal species, we notice that eqn. (21) under bulk conditions, reads

$$1 = 1 + \sum_{j=1}^h s_{ij} f_j^* \quad 1 \leq j \leq h \quad (39)$$

once (11), (30) and (38) have been taken into account. The previous system of equations (39) is homogeneous leading to the trivial solution

$$f_j^* = 0 \quad 1 \leq j \leq h \quad (40)$$

which provides the boundary conditions to prescribe.

The parameters defining the problem

The response given through $\text{grad}(c_0)$ (but not through ϕ) corresponding to equations (3)-(4) with boundary conditions (36)-(40) could be defined with 3 lists of parameters: $\kappa_{a,i}$, $\kappa_{d,i}$ and $\varepsilon_i K_i'$, apart from geometrical factors linked to the boundary conditions such as those related to the diffusion domain not included in ℓ . Recalling the relationship between the association and dissociation parameters given by (9), we notice that only 2 (of the original 3) lists are necessary.

2.- One-dimensional geometries: the degree of lability

As mentioned in the introduction, the concept of lability is related to the availability of the active species arising from the dissociating complexes. In order to qualitatively assess such contribution, the concept of degree of lability for the particular case of only one complex under diffusion limited conditions in spherical geometry was introduced^[18]. The degree of lability for that case was shown to be directly related to the normalised concentration of complex at the active surface:

$$\xi = 1 - c_{ML}^0 / c_{ML}^* \quad (41)$$

where superscript 0 indicates the concentration at the active surface. Indeed, the normalised concentration of the complex species at the active surface is a good measure of the capacity of the complex to dissociate when c_M is depleted and thus of the lability of this complex (for planar geometry results, see Figs 1, 2 and 3 in ref. ^[19]).

The gradient of metal concentration at the origin for this spherical case with only one complex could have been written as:

$$\text{grad } c_M = \frac{c_M^*}{r_0} (1 + \varepsilon K' \xi) \quad (42)$$

We now proceed to generalise this concept to a mixture of ligands in one-dimensional geometries. The first step consists in recognising the key role of the fraction of each complex normalised concentration at the active surface on the flux (and, so, on ϕ).

Adding up the h eqns. in (12) with eqn. (13), one obtains,

$$\nabla^2 \left(\sum_{i=0}^h c_i \right) = 0 \quad (43)$$

As shown in appendix A, restriction to any one dimensional geometry that sustains steady state (such as planar, cylindrical or spherical diffusion in a finite domain or

spherical diffusion in a semi-infinite domain) allows a simple solution of eqn. (43) and the flux to be written in terms of the differences of complex concentrations between the

bulk and the active surface $\sum_{i=0}^h (c_i^* - c_i^0)$. From the definitions, these differences can be

re-organised as:

$$\sum_{i=0}^h (c_i^* - c_i^0) = \left(1 - \frac{c_M^0}{c_M^*}\right) \left[1 + \sum_{i=1}^h \varepsilon_i K_i' \frac{\left(1 - \frac{c_{ML_i}^0}{c_{ML_i}^*}\right)}{\left(1 - \frac{c_M^0}{c_M^*}\right)} \right] \quad (44)$$

It can be shown (see eqns. (A-6)-(A-8)) that

$$\text{grad } c_0 \propto \left[1 + \sum_{i=1}^h \varepsilon_i K_i' \frac{\left(1 - \frac{c_{ML_i}^0}{c_{ML_i}^*}\right)}{\left(1 - \frac{c_M^0}{c_M^*}\right)} \right] \quad (45)$$

By comparing eqn. (45) with (42), we suggest the extension of the concept of degree of lability ^[18;25] to *each* complex of the mixture under any one-dimensional geometry and any steady state regime (i.e. not necessarily diffusion limited conditions) as:

$$\xi_i \equiv \left(1 - \frac{c_{ML_i}^0}{c_{ML_i}^*}\right) / \left(1 - \frac{c_M^0}{c_M^*}\right) \quad (46)$$

However, for simplicity, from now on we only consider here the diffusion limited case (i.e. $c_M^0=0$).

In order to illustrate the suggested ideas, we consider a model mixture whose parameters are gathered in Table 1. There are two complexes: one (ML_1 , a macromolecular species) with strong stability constant, low abundance and low mobility and the other (ML_2 , formed with a small ligand) with weak stability constant, large abundance and same mobility as the metal ion. For these conditions, with a spherical

active surface of radius 10 μm , the degree of lability of the complexes (computed with expressions given below in section 3) is $\xi_1=0.06$ and $\xi_2=0.33$ (see Table 2), as indicated by the height of the corresponding bars in Fig 3.

Using (45) and (46) in eqn. (14), the normalised flux (under diffusion limited conditions) for any one-dimensional geometry can be written

$$\phi = \frac{1 + \sum_{i=1}^h \varepsilon_i K_i' \xi_i}{1 + \sum_{i=1}^h K_i'} \quad (47)$$

To apply this expression numerically one still needs to solve the problem as outlined in section 1, but expression (47) (or (45)) indicates that the potential contribution of each complex ($\varepsilon_i K_i'$) to the overall flux is modulated by the value (relative to the bulk) of its concentration at the active surface (see definition (46)). Indeed, if we plot a bar for each complex i at the abscissa $\varepsilon_i K_i'$ with a height given by the degree of lability for this complex, then the area of the rectangle determined by this bar and the axes (see different fillings of areas in Fig 3) yields the contribution of the chosen complex to the flux indicated by eqn. (47). In order to reproduce the leading 1 in the numerator of eqn (47), a virtual contribution of the free metal can also be included in the figure by adding a bar of unit height at $\varepsilon_i K_i' = 1$.

From Fig 3 (see Table 2 for the numerical values), one concludes that the contribution to the flux of complex ML_1 is much larger than the contribution of ML_2 , despite ML_1 has a lower degree of lability. In this example, the contribution of the free metal to the

overall flux (equivalent to the area between the bar at $\varepsilon_i K_i' = 1$ and the ordinate axis) is negligible in comparison with the contribution of the complex.

A further conclusion from eqn. (47) is that the potential contribution of a particular complex depends on $\varepsilon_i K_i'$. If $\varepsilon_i K_i' \ll 1$, the presence (and lability) of the complex i has a negligible impact on the total flux because the metal has a larger weight, so the assumption $\varepsilon_i K_i' \gg 1$ can be used whenever a simplification is necessary.

Another application of the degree of lability is the possibility of extending the lability criterion concept ^[14;18;25;39;40] to *each* complex in the mixture. One can impose

$$\xi_i > 1/2 \quad (48)$$

as the condition a specific complex ML_i has to fulfil in order to be predominantly labile.

3.- A particular case: semi-infinite spherical diffusion

3.1.- Solution of the problem

In this case a relevant characteristic distance is the electrode radius ($\ell = r_0$). Thus, a normalised co-ordinate can be defined as:

$$\rho \equiv r / \ell = r / r_0 \quad (49)$$

The solution of eqn. (17) with $j=0$ (i.e. $n_j=0$), imposing only the bulk boundary condition (38) and leaving an unknown coefficient (T_0) to be determined from the other boundary condition, is:

$$f_0 = \frac{T_0}{\rho} + 1 \quad (50)$$

The solution of eqn. (17) taking into account the bulk boundary condition (40) can be written (e.g. see Appendix A):

$$f_j = \frac{T_j e^{-\rho\sqrt{n_j}}}{2\rho\sqrt{n_j}} \quad (51)$$

Now, we impose the boundary conditions on the electrode surface ($\rho=1$). In case of diffusion limited flux of M –see (37)-

$$0 = c_0 = \sum_{j=0}^h f_j = T_0 + 1 + \sum_{j=1}^h \frac{T_j e^{-\sqrt{n_j}}}{2\sqrt{n_j}} \quad (52)$$

while the non-flux condition for the complexes -see (36)- becomes

$$\begin{aligned} 0 &= \left(\frac{\partial c_i}{\partial \rho} \right)_{\rho=1} = \sum_{j=0}^h s_{ij} \left(\frac{\partial f_j}{\partial \rho} \right)_{\rho=1} \\ &= -\varepsilon_i K'_i T_0 - \sum_{j=1}^h \left(\frac{\kappa_{a,i}}{\kappa_{d,i} - n_j} \right) \frac{T_j e^{-\sqrt{n_j}}}{2} \left(1 + \frac{1}{\sqrt{n_j}} \right) \end{aligned} \quad (53)$$

Combining eqn.(52) and (53), one obtains

$$1 = \sum_{j=1}^h T_j \frac{e^{-\sqrt{n_j}}}{2} \left(\frac{\kappa_{d,i} + \sqrt{n_j}}{\kappa_{d,i} - n_j} \right) \quad 0 < i \leq h \quad (54)$$

which is a linear system of equations allowing T_1, \dots, T_h to be found. The sought gradient can be computed as

$$\left(\frac{\partial c_0}{\partial \rho} \right)_{\rho=1} = \left(1 + \sum_{i=1}^h \varepsilon_i K'_i \right) \left(\frac{\partial f_0}{\partial \rho} \right)_{\rho=1} = \left(1 + \sum_{i=1}^h \varepsilon_i K'_i \right) \left(1 + \sum_{i=1}^h \frac{T_j e^{-\sqrt{n_j}}}{2\sqrt{n_j}} \right) \quad (55)$$

Thus,

$$\phi = \frac{\left(1 + \sum_{i=1}^h \varepsilon_i K'_i \right)}{\left(1 + \sum_{i=1}^h K'_i \right)} \left(1 + \sum_{i=1}^h \frac{T_j e^{-\sqrt{n_j}}}{2\sqrt{n_j}} \right) \quad (56)$$

After some algebra combining eqns. (10), (21), (32), $n_0=0$, (50), (51), (52) and (9), one finds that the degree of lability for the spherical active surface with semi-infinite diffusion can be computed as:

$$\xi_i = 1 - \sum_{j=1}^h \frac{\sqrt{n_j} T_j e^{-\sqrt{n_j}}}{2(\kappa_{d,i} - n_j)} \quad (57)$$

3.2.- Studying the case of two ligands

With the previous formalism, the expression for the flux found for $h=1$ can be retrieved^[18]. We consider, then, the next case corresponding to $h=2$.

From (25), the general expressions for n_1 and n_2 (regardless of the geometry) are

$$n_j = \left(\kappa_{a,1} + \kappa_{a,2} + \kappa_{d,1} + \kappa_{d,2} \mp \sqrt{(\kappa_{a,1} - \kappa_{a,2} + \kappa_{d,1} - \kappa_{d,2})^2 + 4\kappa_{a,1}\kappa_{a,2}} \right) / 2 \quad (58)$$

where we have chosen the sign minus in front of the root for n_1 and the sign plus for n_2 , so that $n_1 < n_2$ (to be consistent with (26)).

Solving (54) with $h=2$ and replacing into (56), we obtain the general expression for the dimensionless flux:

$$\phi = \frac{1}{1 + K'_1 + K'_2} \frac{\sqrt{n_1} + \sqrt{n_2}}{1 - \frac{(\kappa_{d,1} - 1)(\kappa_{d,2} - 1)}{(1 + \sqrt{n_1})(1 + \sqrt{n_2})} + \frac{\kappa_{d,1}\kappa_{d,2}}{\sqrt{n_1}n_2}} \quad (59)$$

As a first particular case, consider that the complex ML_2 is totally inert ($\kappa_{a,2}$, $\kappa_{d,2}$ and n_1 tend to zero while n_2 tends to $\kappa_{a,1} + \kappa_{d,1}$) while there is no restriction on the lability of ML_1 . Then,

$$\phi_{L2inert} = \frac{1}{1 + K'_1 + K'_2} \left(1 + \frac{\varepsilon_1 K'_1 \sqrt{\kappa_{a,1}}}{\sqrt{\varepsilon_1 K'_1 (1 + \varepsilon_1 K'_1) + \sqrt{\kappa_{a,1}}}} \right) \quad (60)$$

A second particular case consists in the complex ML_2 being fully labile ($\kappa_{a,2}$ and $\kappa_{d,2}$ tend to infinity) while there is no restriction on the lability of ML_1 . Then,

$$\phi_{L2, \text{fully labile}} = \frac{1}{1 + K'_1 + K'_2} \left(1 + \varepsilon_2 K'_2 + \frac{\varepsilon_1 K'_1 \sqrt{\frac{\kappa_{a,1}}{1 + \varepsilon_2 K'_2}}}{\sqrt{\frac{\varepsilon_1 K'_1}{1 + \varepsilon_2 K'_2} \left(1 + \frac{\varepsilon_1 K'_1}{1 + \varepsilon_2 K'_2} \right) + \sqrt{\frac{\kappa_{a,1}}{1 + \varepsilon_2 K'_2}}}} \right) \quad (61)$$

It has been demonstrated ^[18] that decreasing r_0 results in the loss of lability in systems with only one complex, due to diffusion enhancement for smaller r_0 . The same conclusion can be reached for mixtures of ligands, as Fig 4 shows: the normalised current drops when the radius shrinks due to the diminishing of the complex dissociation contributions to the measured flux. In Fig 4 we see a dramatic change in the dynamic behaviour of this particular system for radii around 1 mm. The partial lability of both complexes at this radius can also be seen in the normalised profiles depicted in Fig 5, where we also notice that the reaction layer is much thinner than the diffusion layer ^[19]: the former is seen to disappear around $r/r_0 \approx 1.00002$ in the figure, while the latter extends beyond the plotted region.

For some combination of the parameters (see Fig 6), an almost successive loss of lability of each complex can be seen in the chosen logarithmic representation. This fact indicates a potential use of electrodes with variable radii: the kinetic characteristics of the complexes could be easily retrieved from the succession of recorded steps in the measured fluxes.

For prepared mixtures (i.e. not arising from ligands having a fixed proportion of different types of sites in the same molecule), one has experimental control on

individual $c_{L_i}^*$, which results in a capability of changing K_i' and $\kappa_{a,i}$ which determine a variation in the measured fluxes. Let us assume that we prepare several mixtures with 2 kinds of ligands having each ligand a different complexation site, so that the summation of the concentrations of both sites is fixed in all the mixtures (i.e. $c_{L_1}^* + c_{L_2}^* = \text{const.}$). We caution that experiments done in this way must fulfil the excess of ligand condition for both ligands at each prepared mixture. We see the appearance of a maximum of the flux for a certain intermediate molar fraction $x_1 = c_{L_1}^* / (c_{L_1}^* + c_{L_2}^*)$ in Fig 7, where the kinetic constants of both ligands are not very different, the more mobile ligand is less labile ($\varepsilon_1 > \varepsilon_2$; $k_{d,1} < k_{d,2}$) and we have chosen a suitable radius. From left to right in Fig 7 we increase the proportion of L_1 which diffuses faster than L_2 , but dissociates slower than L_2 . We could interpret the figure saying that adding L_1 increases ϕ at the first additions because of its larger ε_1 , but decreases ϕ at further additions because of its smaller $k_{d,1}$.

We can combine the 2 experimentally controlled variables: r_0 and composition, as seen in Fig 8. The continuous line here is the same as in the previous figure 7, but now looks flatter because of the larger axis scale. If we increase r_0 (dashed line), the complexes are more labile and the limiting step is diffusion: so increasing L_1 yields larger ϕ -values for most of the composition range. If we decrease r_0 (dotted line), diffusion is no longer the limiting step, so the addition of L_1 only means a less labile mixture, and, thus, ϕ decreases.

Conclusions

The coupled system of differential equations (3)-(4) for diffusion-reaction of a mixture of one active species and h inactive complexes can be converted into $h+1$ independent differential equations (17): one Laplace equation (corresponding to the eigenvalue $n_0=0$) and h Helmholtz equations (corresponding to eigenvalues $n_j>0$). This uncoupling (e.g. the set of eigenvalues) is independent of the geometry, but is specially convenient for 1-D geometries, for which the general solutions of Laplace and Helmholtz equations are given in Appendix A. The eigenvalues can be found from (25); being eqn. (58) valid for $h=2$ (with no geometry restriction). All the dynamic information (i.e. mobilities and labilities) about the system can be gathered in two lists of h dimensionless parameters each.

The concept of degree of lability ξ can be extended to any particular complex of the mixture (see eqn. (46)), and thus a lability criterion for each complex can be suggested from eqn. (48). Plotting ξ vs. $\varepsilon K'$ allows the direct visualisation (see Fig. 3) of the contribution of each complex to the global flux.

The flux for spherical geometry and any number of ligands can be obtained finding T_j in the system of linear equations (54) and applying (56). For the case of just 2 ligands, equation (59) can be used straightforwardly. Eqns. (60) and (61) correspond to the particular cases of one of the complexes being totally inert or fully labile.

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Appendix A: The flux depends on the differences between bulk and active surface concentrations.

For any one dimensional geometry, equation (43) can be written ^[41]

$$\frac{d^2 \left(\sum_{i=0}^h c_i \right)}{d\rho^2} + \frac{2\gamma}{\rho} \frac{d \left(\sum_{i=0}^h c_i \right)}{d\rho} = 0 \quad (\text{A-1})$$

where γ is a convergence parameter taking the value 1, $\frac{1}{2}$ or 0 for spherical, cylindrical or planar geometry respectively and ρ is the distance from the centre of symmetry of active surface normalised with respect to the characteristic distance ℓ . The solution of eqn. (A-1) is ^[41]:

$$\sum_{i=0}^h c_i = B_2 + B_1 \rho^{1-2\gamma} \quad \text{if } \gamma \neq 1/2 \quad (\text{A-2})$$

$$\sum_{i=0}^h c_i = B_2 + B_1 \ln \rho \quad \text{if } \gamma = 1/2 \quad (\text{A-3})$$

where the constants B_1 and B_2 are to be determined from the boundary conditions. We prescribe fixed concentrations at the active surface ρ_0 :

$$c_i = c_i^0 \quad 0 \leq i \leq h; \quad \rho = \rho_0 \begin{cases} \rho_0 = 1 & \text{if } \gamma \neq 0 \\ \rho_0 = 0 & \text{if } \gamma = 0 \end{cases} \quad (\text{A-4})$$

and at a certain distance ($\rho = \rho^*$, finite or infinite)

$$c_i = c_i^* \quad 0 \leq i \leq h \quad \rho = \rho^* \quad (\text{A-5})$$

because we want to formulate the solution in terms of the differences of those concentrations (even though the values of c_i^0 are *a priori* unknown).

With these boundaries conditions –and recalling the absence of gradient for the complexes as expressed by (33)- one computes the gradient of c_0 at the active surface as:

$$\left(\frac{d \left(\sum_{i=0}^h c_i \right)}{d \rho} \right)_{\rho=0} = \text{grad } c_0 = \frac{\sum_{i=0}^h (c_i^* - c_i^0)}{\rho^*} \quad \text{if } \gamma = 0 \quad (\text{A-6})$$

$$\left(\frac{d \left(\sum_{i=0}^h c_i \right)}{d \rho} \right)_{\rho=1} = \text{grad } c_0 = \frac{\sum_{i=0}^h (c_i^* - c_i^0)}{\ln \rho^*} \quad \text{if } \gamma = 1/2 \quad (\text{A-7})$$

$$\left(\frac{d \left(\sum_{i=0}^h c_i \right)}{d \rho} \right)_{\rho=1} = \text{grad } c_0 = \frac{\sum_{i=0}^h (c_i^* - c_i^0)}{1 - 1/\rho^*} \quad \text{if } \gamma = 1 \quad (\text{A-8})$$

which mean that the gradients of M are proportional to the differences between bulk and surface concentrations of all species.

Additionally, as eqn. (34) can be re-written as

$$f_0 = \sum_{i=0}^h c_i / \left(1 + \sum_{i=1}^h \varepsilon_i K_i' \right) \quad (\text{A-9})$$

the solution of the Laplace eqn. (17) for $i=0$ ($n_0=0$), i.e. f_0 , can be obtained by simply

dividing eqn (A-2) (or (A-3)) by $1 + \sum_{i=1}^h \varepsilon_i K_i'$.

On the other hand the general solution for the 1-D Helmholtz equation (17) with boundary conditions (40) at $\rho = \rho^*$ and

$$f_j = f_j^0 \quad 1 \leq j \leq h; \quad \rho = \rho_0 \begin{cases} \rho_0 = 1 & \text{if } \gamma \neq 0 \\ \rho_0 = 0 & \text{if } \gamma = 0 \end{cases} \quad (\text{A-10})$$

can be written

$$f_j = \rho^{1/2-\gamma} \left[B_1 I_{1/2-\gamma}(\sqrt{n_j} \rho) + B_2 K_{1/2-\gamma}(\sqrt{n_j} \rho) \right] \quad (\text{A-11})$$

where I and K are first and second kind modified Bessel functions.

Then, the gradients can be found to be:

$$\left(\frac{df_j}{d\rho}\right)_{\rho=0} = -\sqrt{n_j} f_j^0 \coth\left[\sqrt{n_j}\right] \quad \text{if } \gamma = 0 \quad (\text{A-12})$$

$$\left(\frac{df_j}{d\rho}\right)_{\rho=1} = \sqrt{n_j} f_j^0 \frac{I_1(\sqrt{n_j})K_0(\sqrt{n_j}\rho^*) + I_0(\sqrt{n_j}\rho^*)K_1(\sqrt{n_j})}{I_0(\sqrt{n_j})K_0(\sqrt{n_j}\rho^*) - I_1(\sqrt{n_j}\rho^*)K_0(\sqrt{n_j})} \quad \text{if } \gamma = \frac{1}{2} \quad (\text{A-13})$$

$$\left(\frac{df_j}{d\rho}\right)_{\rho=1} = -\sqrt{n_j} f_j^0 \left(\frac{1}{\sqrt{n_j}} + \coth\left[\sqrt{n_j}(\rho^* - 1)\right] \right) \quad \text{if } \gamma = 1 \quad (\text{A-14})$$

APPENDIX B: NOTATION

LATIN SYMBOLS

Symbol	Name	Units	Equation
A	active surface area	m^2	(14)
\mathbf{A}	coupled system matrix	none	(16)
B_1, B_2	constants in the solution of Laplace or Helmholtz equation for one-dimensional geometries	none	(A-2), (A-11)
$c_{L_i}^*$	concentration of ligand L_i (assumed constant)	mol m^{-3}	(2)
c_M	active species concentration	mol m^{-3}	(2)
c_M^*	bulk concentration of the active species (M).	mol m^{-3}	(2)
c_M^0	concentration at the active surface $c_M(r_0)$	mol m^{-3}	(44)
$c_0 \equiv c_M / c_M^*$	normalised concentration of species M	none	(11)
$c_i \equiv \frac{D_{ML_i} K_i c_{ML_i}}{D_M c_{ML_i}^*}$	normalised concentration of species ML_i	none	(10)

\vec{c}	vector of normalised concentrations c_j	none	(20)
D_M	diffusion coefficient of species M	m^2s^{-1}	(4)
D_{ML_i}	diffusion coefficient of species ML_i	m^2s^{-1}	(3)
\mathbf{D}	diagonal matrix containing eigenvalues n_j	none	(18)
f_j	formal species j for a system without coupling	none	(17)
\vec{f}	vector of formal species concentrations f_j	none	(20)
\mathbf{I}	identity matrix	none	(22)
K_i	equilibrium constant for complexation of complex ML_i	$mol^{-1} m^3$	(2)
$K'_i \equiv K_i c_{L_i}^*$	excess-ligand equilibrium constant	none	(2)
$k_{a,i}$ $k_{d,i}$	association and dissociation rate constants	$mol^{-1} m^3 s^{-1}$ and s^{-1}	(1)
ℓ	characteristic length of the system	m	(6)
n_j	eigenvalue for formal species j	none	(17)
r	radial co-ordinate	m	(49)
r_0	radius of the spherical active surface	m	(49), Fig 1
\mathbf{S}	matrix transforming c_j into formal species f_j	none	(20), (32)
\vec{s}_j	column of matrix \mathbf{S}	none	(22), (29)
T_j	constants in the solution of the uncoupled differential equations	none	(50), (51)
t	continuous variable to find eigenvalues	none	(22), (27)

GREEK SYMBOLS

γ	convergence parameter	none	(A-1)
$\varepsilon_i \equiv D_{ML_i} / D_M$	normalised diffusion coefficient for species ML_i	none	(5)
$\kappa_{a,i} \equiv k_{a,i} c_{L_i}^* \ell^2 / D_M$	dimensionless complex formation rate constant for species ML_i	none	(6)
$\kappa_{d,i} \equiv k_{d,i} \ell^2 / D_{ML_i}$	dimensionless complex dissociation rate constant for species ML_i	none	(8)
$\kappa_0 \equiv \sum_{i=1}^h \kappa_{a,i}$	summation of dimensionless association rate constants	none	(7)
ξ_i	degree of lability of complex ML_i	none	(46)
$\rho \equiv r / \ell$	dimensionless co-ordinate	none	(49), (A-1)
ρ^*	dimensionless co-ordinate position where bulk conditions are restored	none	(A-5)
ρ_0	dimensionless co-ordinate of the active surface	none	(A-4)
ϕ	normalised flux	none	(14)

Figure Captions

Fig 1 : Outline of the multicomponent diffusion and reaction (see scheme (1)) towards a surface, where M is the only active species.

Fig 2: Graphical solution of eqn. (27) showing that all the eigenvalues are non-negative. The dotted line stands for the r.h.s of eqn. (27), while the continuous line stands for its l.h.s. Dashed lines stand for the asymptotes.

Fig 3 : Plot of the degree of lability ξ_i vs. the product $\varepsilon_i K_i'$ for each complex (vertical bars marked ML₁ and ML₂). The area of each rectangle delimited by the corresponding vertical bar and the Y-axis (indicated with a different filling for ML₁ and for ML₂) shows the contribution of this complex to the total flux according to eqn. (47). Parameters from Table 1 and $r_0=10 \mu\text{m}$. For completeness purposes, the contribution of the metal (in this case negligible area where there is no room for filling) is also included.

Fig 4 : Impact of changing the electrode radius on the lability. Continuous line stands for the mixture of two partially labile components (i.e. using eqn. (59)). Parameters: $\varepsilon_1=0.039$, $\varepsilon_2=0.020$, $D_M=7.14\times 10^{-9} \text{ m}^2\text{s}^{-1}$, $c_{L_1}^*=0.075 \text{ M}$, $c_{L_2}^*=0.025 \text{ M}$, $K_1=K_2=1.26\times 10^8 \text{ M}^{-1}$ and $k_{a,2}=k_{a,1}=1.35\times 10^{10} \text{ M}^{-1}\text{s}^{-1}$. Marker \square (square) stands for the case of ML₂ being inert and the normalised flux is given by eqn. (60). Marker \times stands for the flux according to eqn. (61) (assuming ML₂ is fully labile).

Fig 5 : Normalised concentration profiles: $c_M(\rho)/c_M^*$ (continuous line), $c_{ML_1}(\rho)/c_{ML_1}^*$ (dashed line) and $c_{ML_2}(\rho)/c_{ML_2}^*$ (dotted line) for the mixture of Fig 4 under semi-infinite spherical diffusion with $r_0=1$ mm.

Fig 6 : The predominant loss of lability of each of the 2 complexes of the mixture as r_0 decreases (continuous line) leads to an almost two-step wave. Parameters: $k_{a,1}=10^6 \text{ mol}^{-1} \text{ m}^3 \text{ s}^{-1}$; $k_{a,2}=10^{14} \text{ mol}^{-1} \text{ m}^3 \text{ s}^{-1}$; $c_{T,L_1}^* = c_{T,L_2}^* = 1 \text{ mol m}^{-3}$; $D_M=10^{-9} \text{ m}^2\text{s}^{-1}$; $K_1 = 1.4 \times 10^{11} \text{ mol}^{-1} \text{ m}^3$; $K_2 = 10^7 \text{ mol}^{-1} \text{ m}^3$; $\varepsilon_1 = 0.01$ and $\varepsilon_2 = 0.1$. Marker \square (square) stands for the hypothetical case of ML_1 being always inert irrespective of the radius as prescribed by eqn. (60) analogously to Fig 4 .

Fig 7 : Variation of the normalised flux ϕ with the molar fraction - x_1 - of ligand L_1 in a binary mixture leading to the appearance of a maximum. Parameters $k_{a,1}=k_{a,2}=10^2 \text{ mol}^{-1} \text{ m}^3 \text{ s}^{-1}$; $c_{L_1}^* + c_{L_2}^* = 10^{-2} \text{ mol m}^{-3}$; $r_0 = 10^{-4} \text{ m}$; $D_M=10^{-9} \text{ m}^2\text{s}^{-1}$; $\varepsilon_1 = 1$; $\varepsilon_2 = 0.01$; $K_1 = 1.4 \times 10^4 \text{ mol}^{-1} \text{ m}^3$ and $K_2 = 4.98 \times 10^3 \text{ mol}^{-1} \text{ m}^3$

Fig 8 : Combination of the effects of changing radius and composition yielding a shift in the maximum position for ϕ . Parameters as in Fig 7 (continuous line), except $r_0 = 10^{-3} \text{ m}$ (dashed line) or $r_0 = 10^{-5} \text{ m}$ (dotted line).

Table 1:

Parameters corresponding to a model mixture.

	Complex 1	Complex 2
diffusion coefficient	$D_{ML_1} = 1.43 \times 10^{-10} \text{ m}^2 \text{ s}^{-1}$	$D_{ML_2} = 7.14 \times 10^{-9} \text{ m}^2 \text{ s}^{-1}$
normalised diffusion coefficient (relative to a metal with $D_M = 7.14 \times 10^{-9} \text{ m}^2 \text{ s}^{-1}$)	$\varepsilon_1 = 0.02$	$\varepsilon_2 = 1$
concentration of free ligand	$c_{L_1}^* = 0.03 \text{ mol m}^{-3}$	$c_{L_2}^* = 3 \text{ mol m}^{-3}$
stability constant	$K_1 = 10^7 \text{ mol}^{-1} \text{ m}^3$	$K_2 = 125.9 \text{ mol}^{-1} \text{ m}^3$
association constant	$k_{a,1} = 1.35 \times 10^7 \text{ mol}^{-1} \text{ m}^3 \text{ s}^{-1}$	$k_{a,2} = k_{a,1} = 1.35 \times 10^7 \text{ mol}^{-1} \text{ m}^3 \text{ s}^{-1}$

Table 2:

Contribution of the different species to the flux in the model mixture given in Table 1 for a spherical active surface of radius 10 μm .

	Metal	Complex 1	Complex 2
$\varepsilon_i K_i'$	-	6000	377
ξ_i	-	0.06	0.33
term in numerator eqn. (47) or area in Fig. 3	1	384	125
relative contribution of the species to the flux: $\varepsilon_i K_i' \xi_i / \left(1 + \sum_{i=1}^h \varepsilon_i K_i' \xi_i \right)$	0.19%	75.33%	24.47%

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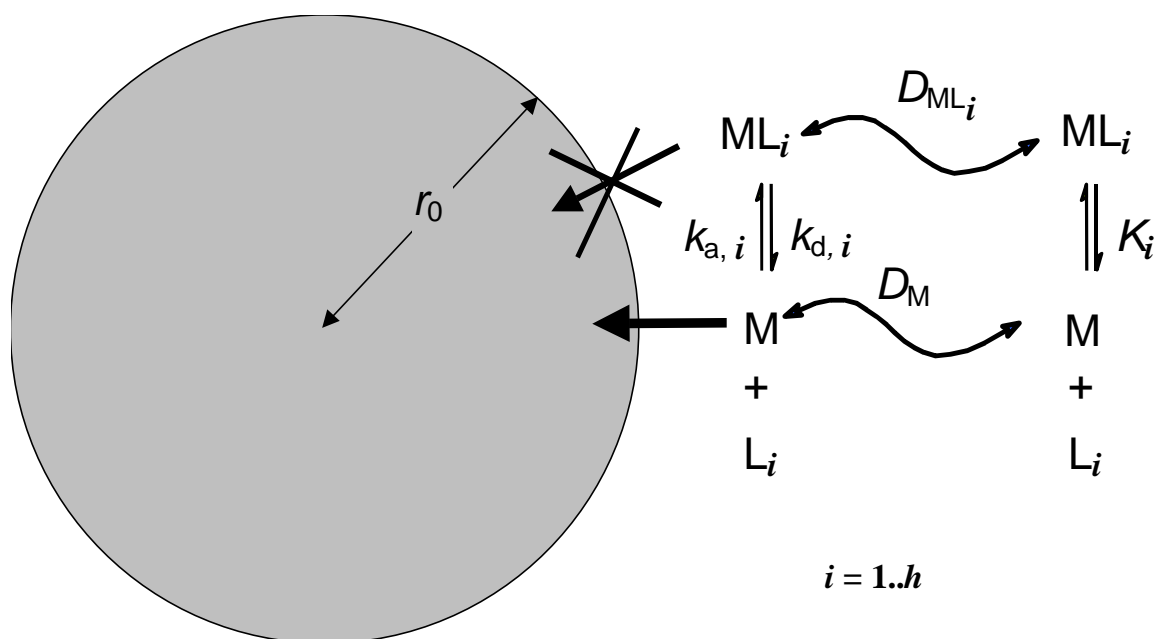


Fig 1

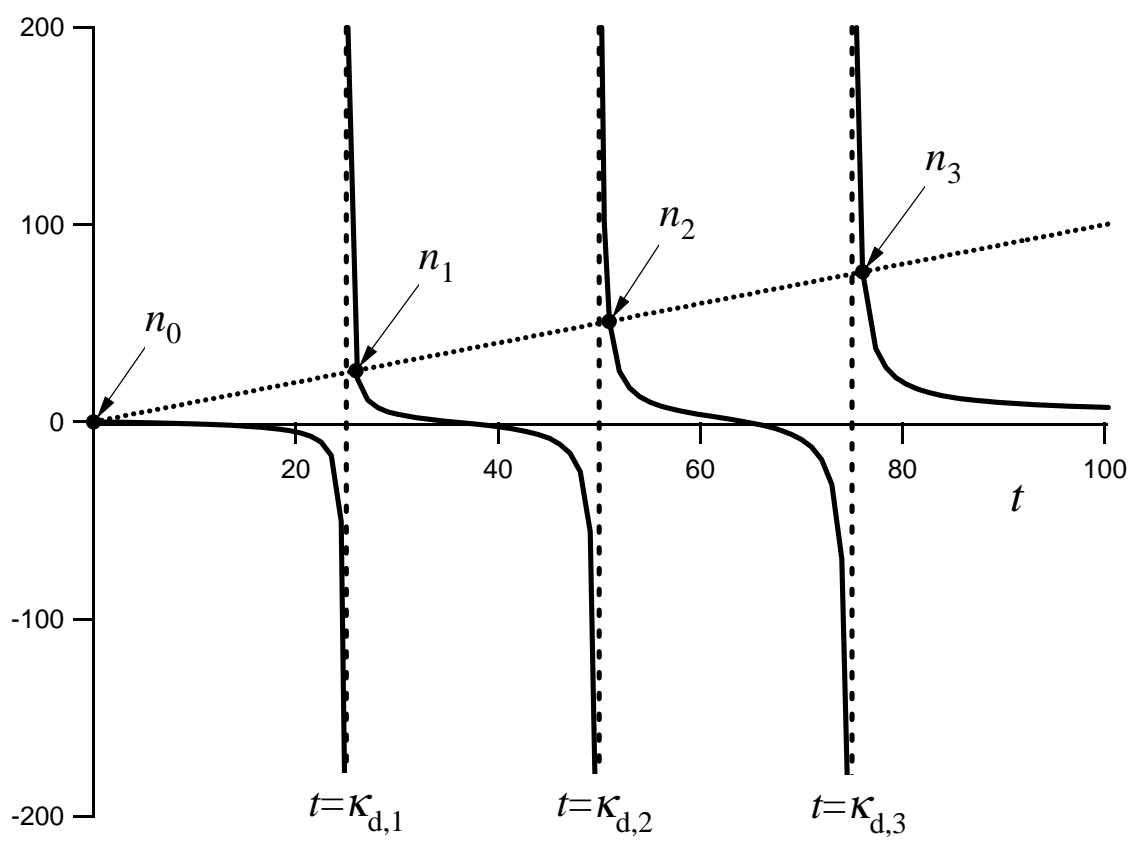


Fig 2

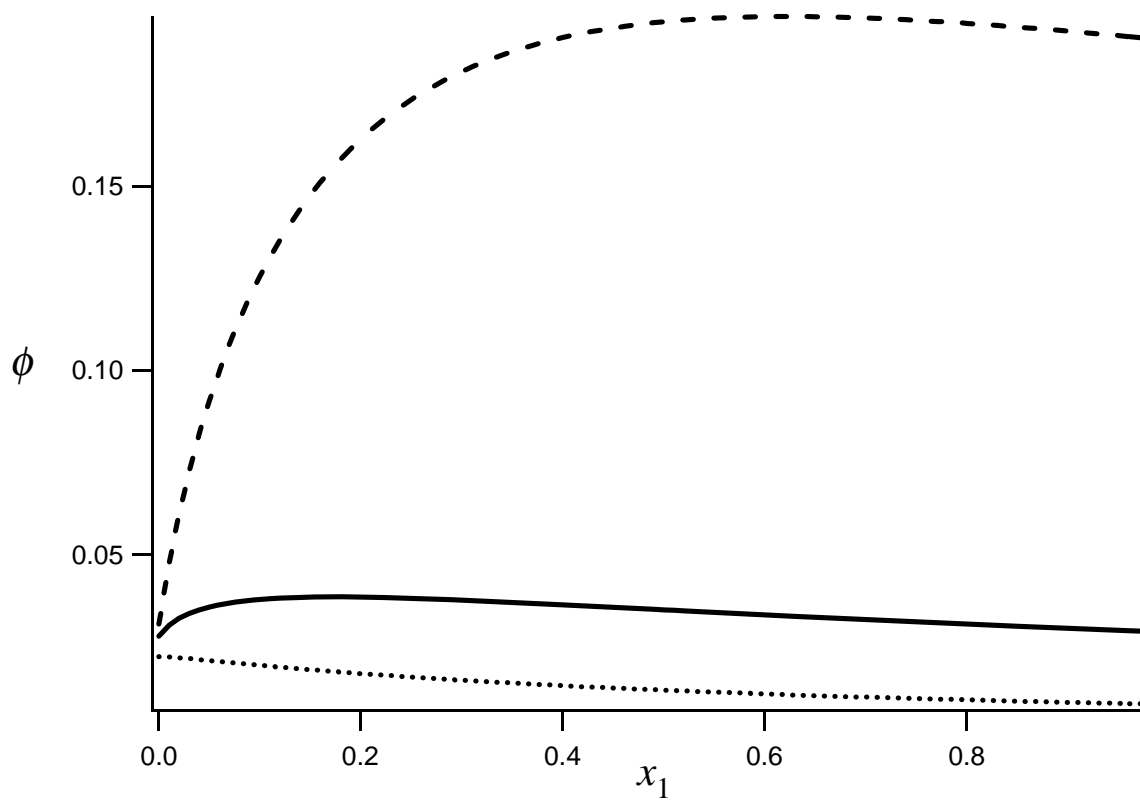


Fig 8